**Supporting information**

**Experimental Methods:** Ferrihydrite naturally exists as a hydrated mineral [5], thus it was important to prepare the mineral such that its hydration was preserved, as opposed to adding known masses of a dried sample. Thus, a ferrihydrite slurry was prepared by adding 0.4 M NaOH to a 50 mM ferric chloride hexahydrate solution, and then centrifuging and washing the resulting precipitate 3 times with 18 MΩ water. The iron content of the prepared ferrhydrite slurry was determined through colorimetry (ferrozine method [6]) and slurry aliquots were added directly to the experiments to achieve the desired mineral mass.

Goethite was prepared through slow air-oxidation of a 50 mM Fe2+ and 100 mM bicarbonate solution over 2 days, and then centrifuged and washed 3 times with 18 MΩ water. The resultant slurry was then oven dried at 70° C for 2 hours, resulting in a mineral powder that was gently ground with a mortar and pestle to homogenize the sample. This dried sample was added directly to experimental bottles.

There are well established methods for preparing and homogenizing natural clay mineral samples for use in experimental work [7]. The montmorillonite ordered from the clay’s society was originally a calcium montmorillonite, but was converted to sodium montmorillonite for easier comparison to other studies. This was achieved by first dispersing the clay with 1 M NaCl, then siphoning the suspended <0.2 µm clay fraction by successively centrifuging and suspending the clay 8 times in DI water. The suspension was flocculated with saturated NaCl, an excess solution removed via centrifugation. The clay suspension was then treated with a 1 M Sodium Acetate solution (pH 5) to remove residual carbonate minerals. The resulting Na-equilibrated montomorillonite was then centrifuged and equilibrated with the experimental background solution, centrifuged again, dried at 70 C overnight, and then gently powdered using mortar and pestle.

After the described mineral preparations here, and in the main text, all minerals’ surface area was measured using a BET surface area analyzer (SI-Table 1).

Displacement of H+ from mineral surfaces can cause sorption experiments to drift in pH over time, requiring some pH adjustment. Following the equilibration period (24 hours), pH was checked and re-titrated to the desired value if necessary; if the pH deviated more than 0.1 pH units, the bottle was allowed to re-equilibrate for 15 minutes after titration, and the re-titration process repeated. This process was sufficient to maintain the experimental pHs described.

**Analytical Methods:** Background concentrations Ra were measured on the scintillation counter to develop a limit of blank of 1.4 counts per second (cps), and activities are reported only for samples that exceeded this value by a factor of 1.5. Supernatant samples collected from the ferrihydrite isotherm, pH 9, did not meet this requirement, so solid associated Ra on the filtered ferrihydrite itself was counted using gamma spectroscopy. A Canberra low energy germanium detector and multichannel analyzer was calibrated using a multinuclide standard from Eckert and Ziegler ([www.ezag.com)](http://www.ezag.com)), and Ra-226 activities were determined using Canberra Genie software using the 186 keV peak. The solid samples on PES filters were placed in scintillation vials, and counted directly on the counter, with the resulting counts being adjusted for ferrihydrite loss during filtration. Gamma spectroscopy was also used to quantify and confirm the 226-Ra standard curve used in scintillation counting.

**Table 1: Mineral BET Surface Areas**

|  |  |
| --- | --- |
| **Mineral** | **Surface Area (m2/g)** |
| Ferrihydrite | 382.9 |
| Goethite | 146.46 |
| Sodium Montmorillonite | 50.162 |
| Pyrite | 0.0685 |

**Surface Complexation Modeling:** Experimental sorption data was fit only by varying radium sorption reaction constants and site densities, preferring literature values for fitted parameters as long as fits were reasonable [8], [9]. Surface area, while a fittable parameter in the models, was not varied, instead using the surface areas reported in Table 1. Solution complexation behavior was accounted for using the SIT database, which includes radium carbonate, sulfate, chloride, and hydroxide complexes, albeit these solution complexes had little impact over the experimental conditions considered.

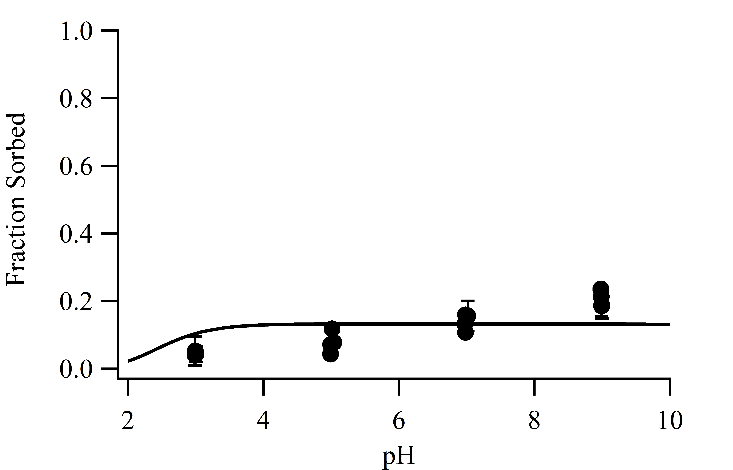
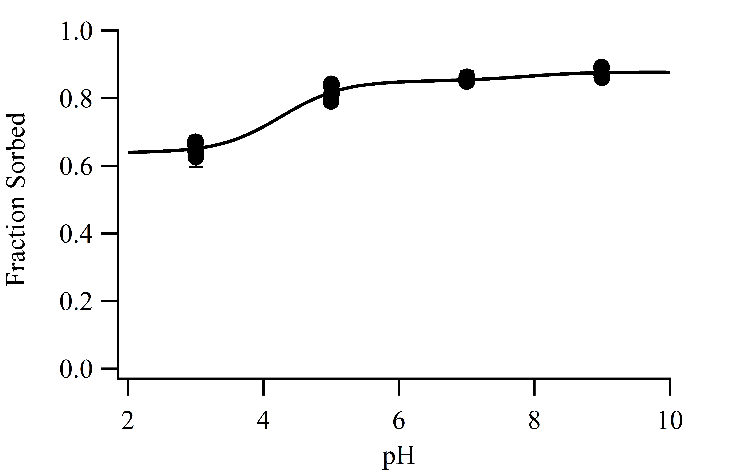
**Sorption Experiment Comparisons:** The literature contains a wealth of studies studying sorption through the collection of experimental isotherms. Comparison of the results in this work with those was easiest done by comparing the measured K­d, compiled in SI-table 2. The wide variety of experimental conditions can make it difficult to make these comparisons, further highlighting the need for a more mechanistic description of Ra sorption, as is provided by SCM.

**Table 2: Comparison of Literature Ra Sorption Experiments and Fitted Kd Values**

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| Mineral | Solid/Solution Ratio (mg/L) | Background Solution | pH | Kd (mL/g) | Source |
| Ferrihydrite | 300  300  25000  10000 | 10 mM NaCl  10 mM NaCl  Seawater  100 mM NaClO4 | 7  9  8.25  7 | 2487  115900  1535  1440 | Experimental  Experimental  [1]  [9] |
| Goethite | 300  300  300  300  25000  10000  500000  500000 | 10 mM NaCl  10 mM NaCl  10 mM NaCl  10 mM NaCl  Seawater  100 mM NaClO4  “pH 1 solution”  “pH 10 solution” | 3  5  7  9  8.25  7  1  10.1 | 0  302.7  573.6  11670  20  50.6  0.752  544 | Experimental  Experimental  Experimental  Experimental  [1]  [9]  [10]  [10] |
| Lepidocrocite | 25000 | Seawater | 8.25 | 174 | [1] |
| Hematite | 25000 | Seawater | 8.25 | 75 | [1] |
| Sodium Montmorillonite | 300  300  3333  50000 | 10 mM NaCl  10 mM NaCl  10 mM NaCl  10 mM NaCl | 5  7  5.25  6.5 | 17750  21470  9700  3724 | Experimental  Experimental  [11]  [12] |

**Pyrite surface treatment:** It is difficult to explain why previous studies have not observed adsorption of Sr to a cleaned and unoxidized pyrite surface, yet we have observed appreciable sorption of Ra to pyrite. Aside from the previously referenced Sr sorption modeling work, we could not find any other study that specifically examined Ra or other group II cation sorption to unoxidized pyrite surfaces. However, many have used a variety of experimental and spectroscopic techniques to study redox active metal ions sorption to the pyrite surface, which often includes redox reactions between the metal ion and pyrite surface groups [13]–[19]. The results of these studies imply that redox active metal ions in solutions with Ra may alter the pyrite surface, consequentially blocking or enhancing Ra sorption. A notable study of an oxidized synthetic pyrite under an extremely O2 free atmosphere (O2 < 0.01 ppm) found that As(III) sorption was similar to other studies of As(III) sorption to unoxidized pyrite in less rigorously anaerobic systems (O2 < 1 ppm) [20]. This suggests that pyrite typically considered “unoxidized” (including the pyrite surface used here) have some differences in surface chemistry compared to a “pristine” synthetic pyrite surface. These differences in pyrite treatment and cleaning during experiments could be the driver for differences in surface sites that result in this unexpected Ra sorption. Two different natural pyrites were used between the experiment here and for studying Sr sorption, thus it is also possible there are also structural differences relating to impurity lattice substitutions or other inclusions into the pyrite [13]. Only spectroscopic investigation of Ra sorption to pyrite surface, however, can elucidate the large differences between Sr and Ra sorption as well as enable the development of an SCM that accurately predicts Ra speciation in the presence of an unoxidized pyrite surface. Indeed the reactivity of the pyrite surface suggests that care must be taken to determine which surface conditions best match those found in the environment.

**Surface Complexation Modeling:**

 Figure SI-1: SCM results omitted from the main text. Lines indicate model fit and points experimental data. Left: SCM of Ra sorption to Na Montmorillonite, using a 2 site reaction formulation and an exchange reaction. Right: SCM of Ra to anaerobic pyrite.

In this work, we report fitting of SCM using established reaction formulations informed by spectroscopic or *ab initio* modeling, which should be more accurate descriptions of solute sorption to a surface [21], [22] (Figure 2, Figure SI-1). There are, however, SCM formulations that are simpler, yet still provide some value; these models pose only the existence of the reaction, making no assumptions about the specific surface complexes that form, which allows for easy comparison of the relative importance of the different minerals for radium retention [23]–[25]. They also may provide utility in larger scale groundwater modeling efforts, as these simpler formulations will be simple to add to larger scale modeling packages that incorporate a variety of biogeochemical and flow processes.

SI-Table 4 provides the results of fitting the experimental data to these simpler reaction formulations. Ferrihydrite and goethite were fit a single site model based on a previous model of iron (hydr)oxide sorption posed by Dzombak, 1990 (Figure SI-1). While that model typically specifies two sites, a strong site controlling sorption at low solute loading, and a weak site contributing to sorption at high solute loading, the simpler model fit here only one site. This likely corresponds to the “strong” site in the Dzombak model due to the low concentration of Ra used experimentally. While a model containing strong and weak sites was fitted, there was little sensitivity to the weak site, so it was omitted. For sodium montmorillonite, a single site with two reactions was considered, but this fit was visually inferior to the reported two site model. The pyrite model reported here matches the model used in the main text, as both use a simple, single site model that makes no assumptions about surface reactions between Ra and pyrite. Comparing the results of the fitted reaction constants reflect experimental observations using Kd. Na-montmorillonite has the largest surface reaction constant, and the largest K­d, followed by ferrihydrite, goethite, and pyrite.

**Table 3: SCM reaction formulas and fitted constants**

|  |  |  |  |
| --- | --- | --- | --- |
| Reactions | Sites (mol/g) | log K | Source |
| Ferrihydrite  ≡FhyOH + H+ = ≡FhyOH2+  ≡FhyOH = ≡FhyO- + H+  ≡FhyOH + Ra2+ = ≡FhyOHRa2+ | 1.87E-3 | 7.92  -8.93  5.7 | [9]  [23]  [23]  Fitting |
| Goethite  ≡GoeOH + H+ = ≡GoeOH2+  ≡GoeOH = ≡GoeO- + H+  ≡GoeOH + Ra2+ = ≡GoeOHRa2+ | 3.99E-3 | 4.8  -10.4  3.5 | [26]  [26]  [26]  Fitting |
| Sodium Montmorillonite  2 ≡Clay-Na + Ra2+ = ≡Clay2-Ra + 2 Na+  ≡ClayOH + H+ = ClayOH2+  ≡ClayOH = ≡Clay- + H+  ≡ClayOH + Ra2+ = ≡ClayOHRa2+  ≡ClayO- + Ra2+ = ClayORa+ | Surf: 3.33E-9  Exch: 8.43E-4 | 0.15  4.5  -7.9  9.8  10.1 | Fitting, clays.org CEC  Fitting  [8]  [8]  Fitting  Fitting |
| Pyrite  ≡PyrSH = ≡PyrS- + H+  ≡PyrS- + Ra2+ = ≡PyrSRa+ | 2.23E-5 | 6.45  -10.5 | [16]  [16]  Fitting |

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